

# BEST ROLEPLAY OUTFITS



# Roblox Roleplay Checklist

**SJ Ball**



## **Roblox Roleplay Checklist:**

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}_2$  what is the

limiting reagent if 8.52 g are actually formed what is the  
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 calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $Al + O_2 \rightarrow Al_2O_3$  4 Al 3  $O_2 \rightarrow 2 Al_2O_3$  b determine the theoretical yield of  $Al_2O_3$  160.0 g Al  $\times 1$  mol Al  $\times 2$  mol  $Al_2O_3$   $\times 101.964$  g  $Al_2O_3$  302.3 g  $Al_2O_3$  26.982 g Al 4 mol Al 1 mol  $Al_2O_3$  c determine the percent yield  
 yield actual  $\times 100$  260.0 g  $\times 100$  86.01 theoretical 302.3 g

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