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web the first two equations have been balanced for you alpha particle α 2 4 beta particle β gamma ray γ proton p neutron n 0 positron β^+ 0 1 30 1 65 1 2 63 2 15 0 30 0 65 30 9 29 2 2 12 6 15 32 30

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web we can write a balanced nuclear equation for this reaction as let s do a practice example identify the product x by writing a balanced nuclear equation $^{137}_{55}\text{Cs} \rightarrow ^{137}_{56}\text{Ba} + x$ because the mass number is the same on both sides of the equation x must have a

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web identify the missing particle in the following nuclear reaction ${}^1_0\text{n} + {}^{235}_{92}\text{u} \rightarrow {}^1_0\text{n} + {}^{137}_{52}\text{te} + 4\text{ }^2_4\text{he} + 97\text{ }^1_0\text{n}$ ${}^{96}_{40}\text{zr} \rightarrow {}^{97}_{40}\text{zr}$

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web answer 4 correct in the reactants and products must be the same part b what is the value of in the following nuclear reaction ${}^{65}_{29}\text{cu} + {}^1_0\text{n} \rightarrow {}^{65}_{29}\text{cu} + 9\text{ }^1_0\text{n}$ express your answer as an integer hint 1 how to approach the problem use rule 2 the total of the subscripts atomic numbers mathematically this can be written as answer

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