

Kinetics of thermal decomposition of Acetaldehyde

$$k_1[\text{CH}_3\text{CHO}] = k_4[\text{CH}_3]^2$$

$$k_4[\text{CH}_3]^2 = k_1[\text{CH}_3\text{CHO}]$$

$$[\text{CH}_3]^2 = \frac{k_1}{k_4} [\text{CH}_3\text{CHO}]$$

$$[\text{CH}_3] = \left(\frac{k_1}{k_4}\right)^{1/2} [\text{CH}_3\text{CHO}]^{1/2} \quad \text{-----(4)}$$

- Substitute (4) in (1) we get

$$\frac{d[\text{CH}_4]}{dt} = k_2 \left(\frac{k_1}{k_4}\right)^{1/2} [\text{CH}_3\text{CHO}]^{1/2} [\text{CH}_3\text{CHO}]$$

$$\frac{d[\text{CH}_4]}{dt} = k_2 \left(\frac{k_1}{k_4}\right)^{1/2} [\text{CH}_3\text{CHO}]^{3/2}$$

- This is the rate equation for thermal decomposition of CH_3CHO and the overall order of this reaction is **3/2**

Thermochemical Kinetics

Thomas Griffiths



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