

Sport Indeed

Thomas Martidale

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D Keegan



Sport Indeed:

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subsequent Δh values $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$ $\Delta h = 226.7 \text{ kJ}$
 $\text{C}_2\text{H}_2(\text{g}) + 2 \text{ H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta h = 91.1 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$ $\Delta h = -241.8 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 hess's law is written as $\Delta H_{\text{target}} = \Delta H_1 + \Delta H_2 + \Delta H_3$ or $\Delta H_{\text{target}} = \sum \Delta H_{\text{known}}$

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web hess's law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the ΔH values for each step to use hess's law two principles must be understood one if an equation is reversed the sign of the ΔH value is also reversed two if an equation is multiplied by a

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) + \text{OH}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ 2 $\text{NaOH}(\text{s}) + \text{H}(\text{aq}) + \text{Cl}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{Na}(\text{aq}) + \text{Cl}(\text{aq})$ 3