

- ✔ Treat your **HOTEL ROOM** like it is your own **HOME**
- ✔ Know the **LOCAL CUSTOMS** and follow them
- ✔ Find out the real **MONETARY VALUE** of things and adhere to those prices
- ✔ Ask for permission before taking somebody's **PHOTO**
- ✔ Respect local **TRADITIONS** and **RELIGION**
- ✔ Learn some basic words and phrases in the local **LANGUAGE**
- ✔ Support the local shops and **ARTISANS**
- ✔ Report **CHILD** sex tourism

- ❌ Don't throw away or leave behind any **EXPENDABLES** that you brought
- ❌ Don't buy illegal **WILDLIFE** products and **HERITAGE** antiques
- ❌ Don't encourage **BEGGING** by giving money to paupers and **STREET KIDS**, rather donate to a local **NGO**
- ❌ Don't overly display your "first-world" **GADGETS**
- ❌ Don't be condescending or arrogant when **TALKING** to locals
- ❌ Don't **JUDGE** an entire culture based on a single individual
- ❌ Don't agree to **BRIBE** officials
- ❌ Don't **LITTER**, even if locals do



Offset your **CARBON FOOTPRINT**



Minimize air travel in favor of land or sea **TRANSPORT**



Join a local **NGO** or do some **VOLUNTARY** work  
 ↳ learn how you can make a difference



Set a **GOOD EXAMPLE** for locals and fellow travelers  
 ↳ how you behave will influence how locals see future travelers



Curb your **FREEDOM** of **SPEECH** about government and religion  
 ↳ not doing so could get you and your local friends in serious trouble



Stay in family-run hotels/hostels or do a **HOMESTAY**



Carry **REUSABLE CONTAINERS**  
 ↳ use them at street food and drink stalls instead of the disposable ones provided



"Take nothing but pictures. Leave nothing but footprints. Kill nothing but time."

# Sustainable Travel Tutorial

**M Planty**



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web donor whereas a base generically represented by  $B$  is a proton acceptor when acids and bases combine they swap a proton  $H^+$   $B + H^+ \rightarrow HB$  strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $\text{pH} = \log 0.100 = 1.000$  ph  $\log 0.100 = 1.000$  b titrant volume 12.50 ml

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