



# The Creative Arts

**Didier Musso**



## **The Creative Arts:**

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web 1.csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH<sub>3</sub> can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction Al + O<sub>2</sub> → Al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160.0g Al × 1 mol Al / 27g Al × 2 mol Al<sub>2</sub>O<sub>3</sub> / 3 mol Al × 101.96g Al<sub>2</sub>O<sub>3</sub> / 1 mol Al<sub>2</sub>O<sub>3</sub> = 302.3g Al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual / yield theoretical × 100 = 260.0g / 302.3g × 100 = 86.01%

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web 7Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33g of H<sub>2</sub> are left over 13

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