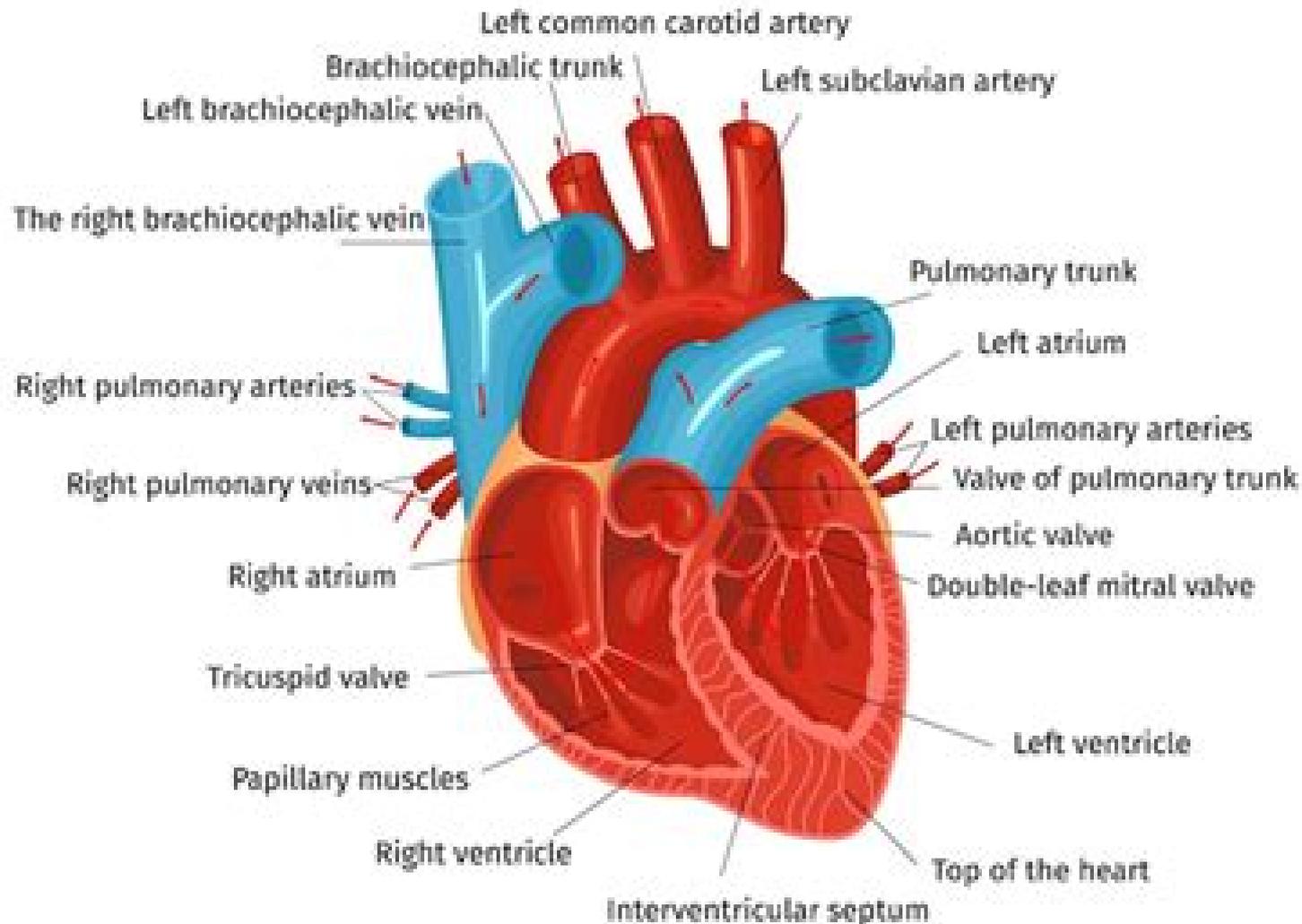


Human Heart Structure and Function



Your Heart And How To

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web limiting reactant theoretical and percentage yields key $4\text{K}_2\text{S} + 2\text{H}_2\text{O} + 4\text{KOH} \rightarrow 3\text{O}_2 + \text{G}$ if a reaction vessel contains 0 mol K_2O and 0 mol H_2O what is the limiting reactant

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g } \text{Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol } \text{Al}_2\text{O}_3}{2 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 298.2 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{260.0 \text{ g}}{298.2 \text{ g}} \times 100 = 87.2\%$

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web 7 Na $2\text{C}_2\text{O}_4$ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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 $80 \text{ g } \text{I}_2\text{O}_5 \cdot 1 \text{ mol } \text{I}_2\text{O}_5 / 1 \text{ mol } \text{I}_2 \cdot 253.8 \text{ g } \text{I}_2\text{O}_5 = 1 \text{ mol } \text{I}_2\text{O}_5 \cdot 28 \text{ g } \text{C} / 1 \text{ mol } \text{C}$

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