

GENERAL ORGANIC CHEMISTRY

संज्ञक पदार्थ

REACTION

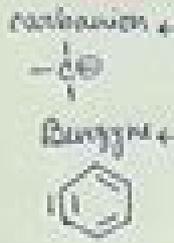
IONIC REACTION

$\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow 2\text{NaCl} + \text{BaSO}_4$
 - no intermediate at all.

MOLECULAR REACTION

- Rem of organic chemistry.
 $\text{CH}_2=\text{CH}=\text{CH}_2 + \text{Cl} \rightarrow \text{CH}_2-\text{CH}-\text{CH}_2$
 - Formation of intermediates occur.

संज्ञक पदार्थ \rightarrow e^- rich



INTERMEDIATE

- very high energy
- highly unstable
- survive for very short time
- In a run only one type of intermediate is formed
- degree can be different
- survived by e^- movement

e^- deficient

- survived by e^- releasing effects
- \rightarrow carbocation $-\text{C}^\oplus$
- \rightarrow free radical $-\dot{\text{C}}$
- \rightarrow carbene $:\text{C}:$
- \rightarrow Nitrene $-\text{N}:$

SUBSTRATE + REAGENT

PRODUCT

INTERMEDIATE

survived by e^- movements

Inductive effect

$+I \rightarrow e^-$ releasing
 $-I \rightarrow e^-$ withdrawing

permanent transfer of σ -Bond e^- due to difference in electronegativity.

Hyper conjugation effect

transfer of $\text{C}-\text{H}$ σ -Bond electrons in electron deficient system.

Resonance

movement of π and lone pair e^- to establish electron e^- equilibrium.

$+R/H$; $-R/-H$

Straining Resonance

due to angle strain.

Electromeric Effect

the temporary transfer of π -bond electrons in the presence of electrophile or nucleophile.

$+E, -E$

→ Highly stable
 → Less reactive
 → never initiates reaction

→ Less stable
 → more reactive
 → initiates rxn

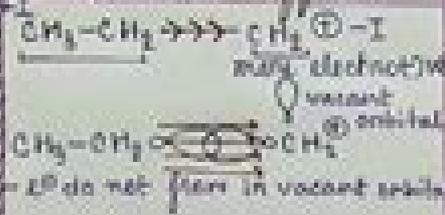
MAJOR PRODUCT

- more than 50%
- from more stable intermediate

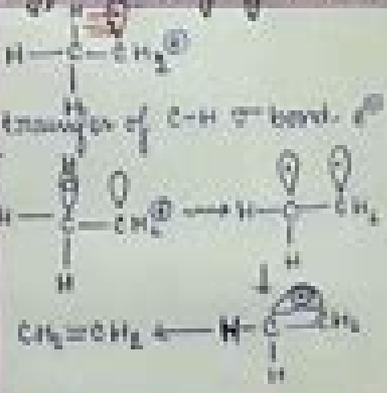
MINOR PRODUCT

- less than 50%
- from less stable intermediate

Inductive Effect



Hyper Conjugation



Understanding Organic Chemistry

Guillaume Favre



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