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The Library Trustee A Practical Guidebook

Sebastian Brünink



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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + \text{Ag}_2\text{O} + 2\text{Cu}_2\text{O} + 2\text{AgNO}_3 + 2\text{Ag}_2\text{O}$$

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27.0 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mole} = 302.3 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{26.982 \text{ g } \text{Al}_2\text{O}_3}{302.3 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 8.93\%$

yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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 $80 \text{ g } i_2O_5 \quad 1 \text{ mol } i_2O_5 \quad 1 \text{ mol } I_2 \quad \times \quad 1 \quad 333.8 \text{ g } i_2O_5 \quad 1 \text{ mol } i_2O_5 \quad 28 \text{ g } CO \quad 1 \text{ mol } CO$

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