



# INSTRUCTIONAL MEDIA



# Using Instructional Media

**Ying Liu**



## **Using Instructional Media:**

## **Using Instructional Media** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such may be the essence of the book **Using Instructional Media**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

<https://ftp.thebrandexperience.com/book/book-search/index.jsp/using%20dollars%20and%20sense%20macintosh%20version.pdf>

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web 1 consider the reaction  $\text{i}_2\text{o}_5 \text{ g } 5 \text{ co g } 5 \text{ co}_2 \text{ g } \text{i}_2 \text{ g } \text{ a } 80 \text{ 0 grams of iodine v oxide } \text{i}_2\text{o}_5 \text{ reacts with } 28 \text{ 0 grams of carbon monoxide co determine the mass of iodine } \text{i}_2 \text{ which could be produced b if in the above situation only } 0 \text{ 160 moles of iodine } \text{i}_2 \text{ was produced i what mass of iodine was produced}$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting\_reagents\_theoretical\_actual\_and\_percent\_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles } \text{O}_2$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   $\frac{160 \text{ g } \text{Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol } \text{Al}_2\text{O}_3}{2 \text{ mol } \text{Al}} \times 101.96 \text{ g/mol } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 = \frac{260 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7  $\text{Na}_2\text{C}_2\text{O}_4$  is the limiting reactant percent yield 86.68% only four molecules can be made 9 this amount cannot be weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of  $\text{H}_2$  are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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