

# Kinetics of thermal decomposition of Acetaldehyde

$$k_1[\text{CH}_3\text{CHO}] = k_4[\text{CH}_3]^2$$

$$k_4[\text{CH}_3]^2 = k_1[\text{CH}_3\text{CHO}]$$

$$[\text{CH}_3]^2 = \frac{k_1}{k_4} [\text{CH}_3\text{CHO}]$$

$$[\text{CH}_3] = \left(\frac{k_1}{k_4}\right)^{1/2} [\text{CH}_3\text{CHO}]^{1/2} \quad \text{-----(4)}$$

- Substitute (4) in (1) we get

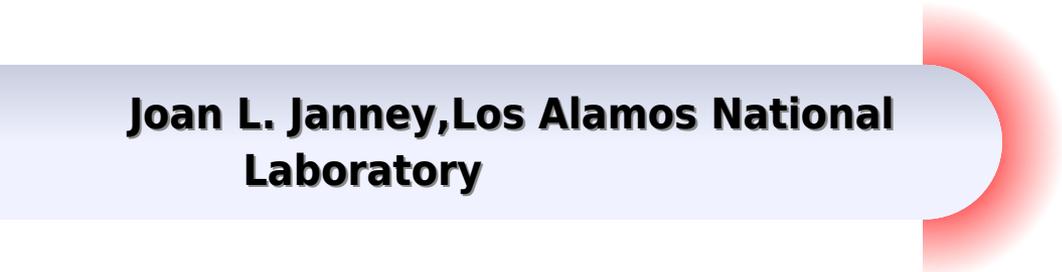
$$\frac{d[\text{CH}_4]}{dt} = k_2 \left(\frac{k_1}{k_4}\right)^{1/2} [\text{CH}_3\text{CHO}]^{1/2} [\text{CH}_3\text{CHO}]$$

$$\frac{d[\text{CH}_4]}{dt} = k_2 \left(\frac{k_1}{k_4}\right)^{1/2} [\text{CH}_3\text{CHO}]^{3/2}$$

- This is the rate equation for thermal decomposition of  $\text{CH}_3\text{CHO}$  and the overall order of this reaction is **3/2**

# Thermochemical Kinetics

**Joan L. Janney, Los Alamos National  
Laboratory**



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**NBS Special Publication** ,1968 *Functional Properties of Advanced Engineering Materials and Biomolecules* Felipe A. La Porta,Carlton A. Taft,2021-05-17 This book shows how a small toolbox of experimental techniques physical chemistry concepts as well as quantum classical mechanics and statistical methods can be used to understand explain and even predict extraordinary applications of these advanced engineering materials and biomolecules It highlights how improving the material foresight by design including the fundamental understanding of their physical and chemical properties can provide new technological levels in the future

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*CHEMICAL KINETICS OF SMALL ORGANIC RADICALS* Zeev B. Alfassi,1988-05-31 Cover title CRC chemical kinetics of small organic radicals

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