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 $2 na + f_2 \rightarrow 2 naf$
 $cl_2 + h_2 \rightarrow 2 hcl$
 $2 h_2 + o_2 \rightarrow 2 h_2o$
 $pb + oh \rightarrow 2 h_2o + 1 pbcl_2$
 $al + br_3 \rightarrow 3 albr_3$
 $k_2so_4 \rightarrow 6 kbr + 1 al_2so_4 + 3 ch_4 + 2 o_2$
 $\rightarrow 1 co_2 + 2 h_2o$
 $c_3h_8 + 5 o_2 \rightarrow 3 co_2 + 4 h_2o$
 $c_8h_{18} + 25 o_2 \rightarrow 16 co_2 + 18 h_2o$

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 solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas
 gaseous butane c_4h_{10} reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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 $1 na + 3 po_4 \rightarrow 3 koh + 3 naoh + 1 k_3po_4 + 1 mgf_2 + 1 li_2co_3 + 1 mgco_3 + 2 lif + 1 p_4 + 3 o_2 + 2 p_2 + 2 o_3 + 2 rbno_3 + 1 bef_2 + 1 be + no_3 + 2 rbf$

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 $2 na + f_2 \rightarrow 2 naf$
 $cl_2 + h_2 \rightarrow 2 hcl$
 $2 h_2 + o_2 \rightarrow 2 h_2o$
 $5 pb + oh \rightarrow 2 h_2o + 1 pbcl_2$
 $6 al + br_3 \rightarrow 3 albr_3$
 $2 k_2so_4 \rightarrow 6 kbr + 1 al_2so_4 + 3 ch_4 + 2 o_2$
 $\rightarrow 1 co_2 + 2 h_2o$
 $8 c_3h_8 + 5 o_2 \rightarrow 3 co_2 + 4$

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 $1 o_2 + 2 nh_3 \rightarrow 3 h_2o + 2 o_2 + 3 h$

2 o 2 h 2 o o 2 4 rb h 2 o rboh h 2 5 cl 2 kbr kcl br 2 6 n 2 h 2 nh 3 7 ndf 3 ca nd caf 2 8 ch 4 h 2 o co h 2 9 c 2 h 5 oh

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