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# Technology Management Society

**R Pring**



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web 1 consider the reaction  $\text{i}_2\text{O}_5 + 5 \text{CO} \rightarrow 5 \text{CO}_2 + \text{I}_2$  a 80.0 grams of iodine v oxide  $\text{i}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced b if in the above situation only 0.160 moles of iodine  $\text{i}_2$  was produced i what mass of iodine was produced

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation  
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$   
 is your theoretical yield so percent yield =  $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$   
 problems 6 & 8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation  
 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$   
 if the reaction of 21.3 grams of O<sub>2</sub> produces a 47.3% yield how many grams of CO would be produced  
 step one figure out how many moles of the limiting reagent you have  
 $21.3 \text{ g O}_2 \times \frac{1 \text{ mole}}{32 \text{ g}} = 0.666 \text{ moles O}_2$   
 $0.666 \text{ moles O}_2 \times \frac{3 \text{ moles CO}}{3 \text{ moles O}_2} = 0.666 \text{ moles CO}$   
 $0.666 \text{ moles CO} \times 28 \text{ g/mole} = 18.6 \text{ g CO}$

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web percent yield is a measure of the efficiency of the experimental design  
 yield efficiency =  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$   
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction  
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$   
 4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub>  
 b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub>  
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$   
 $26.98 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 26.98 \text{ g Al}_2\text{O}_3$   
 c determine the percent yield  
 yield actual =  $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$   
 theoretical = 302.3 g

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H<sub>2</sub> are left over 13

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web learning objectives define and determine theoretical yields actual yields and percent yields identify a limiting reagent from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

**limiting reactant and reaction yields article khan academy** - Jan 08 2023

web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

**8 5 limiting reactant theoretical yield and percent yield** - Dec 07 2022

web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is complete

**12 3 limiting reagent and percent yield worksheet answer key** - Jan 28 2022

web limiting reagent and percent yield practice worksheet answer key limiting reactant yield practice worksheet 1 methanol  $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of CO is reacted with 8.6 g of  $\text{H}_2$

limiting reactant and percent yield worksheet with key - Oct 05 2022

web honors chemistry 1b limit reactant and percent yield worksheet with excess calculation name sto 4 solve stoichiometric problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical reaction

*limiting reactant and percent yield practice hs science* - Jun 01 2022

web limiting reactant and percent yield practice name 1 consider the following reaction  $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{}_3\text{PO}_4 + \text{N}_2\text{O}$  which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

*limiting reagents and percentage yield worksheet answers* - Sep 04 2022

web limiting reagents and percentage yield worksheet 1 consider the reaction  $\text{I}_2\text{O}_5 + \text{CO} \rightarrow \text{I}_2 + \text{CO}_2$  80.0 grams of iodine v oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine  $\text{I}_2$  which could be produced  $80 \text{ g } \text{I}_2\text{O}_5 \cdot 1 \text{ mol } \text{I}_2\text{O}_5 / 1 \text{ mol } \text{I}_2 \cdot 333.8 \text{ g } \text{I}_2\text{O}_5 / 1 \text{ mol } \text{I}_2\text{O}_5 = 28 \text{ g CO} \cdot 1 \text{ mol CO}$

**7 3 limiting reactant and percent yield problems** - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  label 3 7 3