



# Guide Work From Home Setup

**Roman Wölfel**



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web elements in the second row of the periodic table place their electrons in the  $2n$  shell as well as the  $1n$  shell for instance lithium li has three electrons two fill the  $1s$  orbital and the third is placed in the  $2s$  orbital giving an electron configuration of  $1s^2 2s^1$

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web the electron removed during the ionization of beryllium  $he 2s^2$  is an  $s$  electron whereas the electron removed during the ionization of boron  $he 2s^2 2p^1$  is a  $p$  electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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