

The Animal Doctors Answer

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web 1 molarity $m = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$ 4500 g / (82.5 g/mol × 0.129 L) = 4.5 mol / 0.129 L = 34.9 M
2 molality $m = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$ 825 g / (95.79 g/mol × 0.00581158 L) = 1.5 mol / 0.00581158 L = 258.1 M
3 mole fraction $n_2 = \frac{\text{mass}}{\text{molar mass}}$ 4.0 g / 82.5 g/mol = 0.0485 mol
4 0.0485 mol / (0.0485 mol + 0.00581158 mol) = 0.894

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web apr 7 2014 molarity moles of solute / liters of solution
8 4.22 a first convert 250 ml to liters 250 / 1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M
9 c a solution with molarity 2 requires 2 M of NaOH per liter so 4 × 2.8 M = 11.2 M
10 a a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture can be calculated using the following formula $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2}{V_1 + V_2}$ 1.5 × 480 + 1.2 × 520 / (480 + 520) = 1.344 M
q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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2 what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution
3 140 g of KCl is dissolved in 600 ml of water what is the molarity
4 724 g of ammonium phosphate in 4500 ml of alcohol

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web determine the molarity for each of the following solutions
0 444 mol of CO₂ in 0.654 L of solution
98.0 g of phosphoric acid H₃PO₄ in 1.00 L of solution
0.2074 g of calcium hydroxide Ca(OH)₂ in 40.00 ml of solution
10.5 kg

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