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**Victor M. Corman**



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web jan 25 2023 h<sup>+</sup> is called conjugate acid of base oh<sup>-</sup> thus every base has its conjugate acid and every acid has its conjugate base 3 lewis concept g n lewis 1923 proposed the concept of acids and bases upon the electronic theory of valency to include those reactions where no proton transfer takes place according to this concept an acid

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web arrhenius bases an arrhenius base is defined as any species that increases the concentration of hydroxide ions oh<sup>-</sup> in aqueous solution an example of an arrhenius base is the highly soluble sodium hydroxide naoh sodium hydroxide dissociates in water as follows na<sup>+</sup> oh<sup>-</sup> a q na<sup>+</sup> oh<sup>-</sup> a q in water sodium hydroxide

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web work done by svante arrhenius in the late 19th century set the groundwork for our current understanding of acid base

theory despite the relatively limited technology available to him he was able to develop theories that adequately explained the properties and chemical reactivity of many acids and bases

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web the arrhenius theory of acids and bases acids are substances that produce hydrogen ions in solution bases are substances that produce hydroxide ions in solution neutralization happens because hydrogen ions and hydroxide ions react to produce water  $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$   $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{HCl} \text{l}$  hydrochloric

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web aug 15 2020 the equilibrium formulated in scheme 1 between hydrogen ion and the corresponding acid and base can be called a simple acid base equilibrium by mixing two simple systems a double acid base system and an acid base equilibrium result that can always be formulated as follows  $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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web key points a brønsted lowry acid is any species that is capable of donating a proton  $\text{H}^+$  a brønsted lowry base is any species that is capable of accepting a proton which requires a lone pair of electrons to bond to the  $\text{H}^+$  water is amphoteric which means it can act as both a brønsted lowry acid and a brønsted lowry base

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web the  $\text{A}^-$  ion is a base because it is accepting a proton from the  $\text{H}_3\text{O}^+$  the reversible reaction contains two acids and two bases we think of them in pairs called conjugate pairs when the acid  $\text{HA}$  loses a proton it forms a base  $\text{A}^-$  when the base  $\text{A}^-$  accepts a proton back again it obviously reforms the acid  $\text{HA}$

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discovery by 1890 the swedish chemist svante arrhenius 1859 1927 was able to formulate the first useful theory of acids  
arrhenius definition an acidic substance is one whose molecular unit contains at least one

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electrons instead of proton transfer and specifically stated that an acid is a species that accepts

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principle is the acid base behavior of molecules in the general chemistry sequence the brønsted lowry theory is stressed  
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