

Thermodynamic properties of ethane from calorimetric measurements

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The specific heat capacity at constant pressure c_p and the isenthalpic Joule-Thomson coefficient μ of ethane have been determined by flow calorimetry in the temperature range between 283 and 473 K at pressures up to 10 MPa. The experimental results and derived properties are reported and compared with corresponding values determined from experimental p, V, T data and existing equations of state. It is shown that the non-analytical equation of state published recently by Goodwin yields the residual part of the heat capacity within 2 per cent, the residual part of the enthalpy within 0.3 per cent, and isothermal entropy differences within 0.3 per cent.

1. Introduction

Thermodynamic properties of gaseous hydrocarbons are important for various process calculations. In most previous investigations the determination of thermal properties like heat capacity, enthalpy, and entropy, is based on experimental p, V, T data and ideal-gas properties. Since first and second derivatives of the experimental p, V, T data or of the equation of state are needed for the calculation of derived thermal properties, it is very difficult to estimate reliably the errors in the derived properties. Therefore, the direct experimental determination of thermal properties is also of interest.

In this investigation a flow apparatus described previously^(1,2) was used for measuring the isobaric specific heat capacity c_p and the isenthalpic Joule-Thomson coefficient $\mu = (\partial T/\partial p)_H$ of gaseous ethane at temperatures between 283 and 473 K and pressures up to 10 MPa. The investigated range of state includes the region near the critical point (critical temperature $T_c = 305.4$ K; critical pressure $p_c = 4.88$ MPa). From the experimental results the enthalpy, the entropy, and the compression factor were determined. The results are compared with existing experimental data, and with the results obtained from the Benedict-Webb-Rubin equation of state^(3,4) and from a non-analytical equation of state published recently by Goodwin.⁽⁵⁾ The comparison shows that thermal properties derived from the p, V, T measurements by Michels *et al.*⁽⁶⁾ and from the non-analytical equation of state by Goodwin are in excellent agreement with the results from the calorimetric measurements. The BWR-equations give a good representation of the calorimetric results only at higher temperatures.

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Huangqi Zhang



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