

THEORY AND APPLICATION OF CONDUCTIVITY

BACKGROUND

Conductivity is a measure of how well a solution conducts electricity. To carry a current a solution must contain charged particles, or ions. Most conductivity measurements are made in aqueous solutions, and the ions responsible for the conductivity come from electrolytes dissolved in the water. Salts (like sodium chloride and magnesium sulfate), acids (like hydrochloric acid and acetic acid), and bases (like sodium hydroxide and ammonia) are all electrolytes. Although water itself is not an electrolyte, it does have a very small conductivity, implying that at least some ions are present. The ions are hydrogen and hydroxide, and they originate from the dissociation of molecular water. See Figure 1.

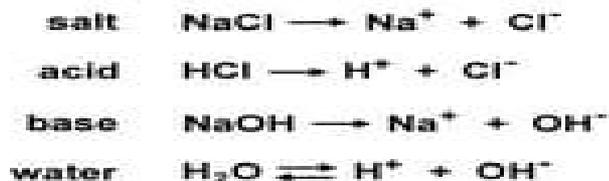


FIGURE 1. Salts, acids, and bases are electrolytes. They dissolve in water to form ions. Although water is not an electrolyte, a very small concentration of hydrogen and hydroxide ions are always present in pure water.

Conductivity is not specific. It measures the total concentration of ions in solution. It cannot distinguish one electrolyte or ion from another.

Not all aqueous solutions have conductivity. Solutions of non-electrolytes, for example sugar or alcohol, have no conductivity because neither sugar nor alcohol contains ions nor do they produce ions when dissolved in water.

APPLICATIONS OF CONDUCTIVITY

Conductivity measurements are widely used in industry. Some important applications are described below.

- **Water treatment.** Raw water as it comes from a lake, river, or the tap is rarely suitable for industrial use. The water contains contaminants, largely ionic, that if not removed will cause scaling and corrosion in plant equipment, particularly in heat exchangers, cooling towers, and boilers. There are many ways to treat water, and different treatments have different goals. Often the goal is demineralization, which is the removal of all or nearly all of the contaminants. In other cases the goal is to remove only certain contaminants, for example hardness ions (calcium and magnesium). Because conductivity is a measure of the total concentration of ions, it is ideal for monitoring demineralizer performance. It is rarely suitable for measuring how well specific ionic contaminants are being removed.

Conductivity is also used to monitor the build up of dissolved ionic solids in evaporative cooling water systems and in boilers. When the conductivity gets too high, indicating a potentially harmful accumulation of solids, a quantity of water is drained out of the system and replaced with water having lower conductivity.

- **Leak detection.** Water used for cooling in heat exchangers and surface condensers usually contains large amounts of dissolved ionic solids. Leakage of the cooling water into the process liquid can result in potentially harmful contamination. Measuring conductivity in the outlet of a heat exchanger or in the condenser hotwell is an easy way of detecting leaks.

- **Clean in place.** In the pharmaceutical and food and beverage industries, piping and vessels are periodically cleaned and sanitized in a procedure called clean-in-place (CIP). Conductivity is used to monitor both the concentration of the CIP solution, typically sodium hydroxide, and the completeness of the rinse.

- **Interface detection.** If two liquids have appreciably different conductivity, a conductivity sensor can detect the interface between them. Interface detection is important in a variety of industries including chemical processing and food and beverage manufacturing.

- **Desalination.** Drinking water desalination plants, both thermal (evaporative) and membrane (reverse osmosis), make extensive use of conductivity to monitor how completely dissolved ionic solids are being removed from the brackish raw water.

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