

Zion's Praises

Sing Unto the Lord

Zions Praises

Ian Pickup



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web limiting reactant theoretical and percentage yields key $4 KO_2 + 2 H_2O \rightarrow 4 KOH + 3 O_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $137.1 \text{ g } \text{H}_2\text{O} / 149.4 \text{ g } \text{H}_2\text{O} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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 $21.3 \text{ g } \text{O}_2 / 32 \text{ g/mole } = 0.666 \text{ moles } \text{O}_2$
 $0.666 \text{ moles } \text{O}_2 \times 3 \text{ g } \text{CO} / \text{mole } \text{O}_2 = 1.998 \text{ g } \text{CO}$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained / calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
 b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 / \text{mol } \text{Al}_2\text{O}_3 = 150.0 \text{ g } \text{Al}_2\text{O}_3$
 c determine the percent yield
 $\text{yield} = \frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{150.0 \text{ g}} \times 100 = 173.3\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4 33 g of h₂ are left over 13

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