

# Cognitive Therapy Techniques

A Practitioner's Guide

# Ebook Therapy Techniques

**Sewon Kang**



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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at stp  $V_1/n_1 = V_2/n_2$   $0.0896 \text{ l} / n_2 = 22.4 \text{ l} / 1 \text{ mole}$   $n_2 = 0.0040 \text{ moles}$  12 experimental molar mass g mole of butane 0.23 g / 0.0040 moles = 58 g mole 13 the molecular formula of butane is  $C_4H_{10}$

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