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# The Climate Of Hell

**Didier Musso**



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + 2 \text{ C}_2\text{H}_2(\text{g}) + 5 \text{ O}_2(\text{g}) + 2 \text{ C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \rightarrow \Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \rightarrow \Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \rightarrow \Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow \Delta h = 285 \text{ kJ/mole}$

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web this activity provides a demonstration of hess law using three reactions the solubility  $\text{NaOH}$  in water the solubility  $\text{NaOH}$  in  $\text{HCl}$  and the reaction of a solution of  $\text{HCl}$  and a solution of  $\text{NaOH}$

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multiply the mass by the change in temperature and the  $c$  given trial 1 use Hess's law and the following equations and  $\Delta H$  values to determine the heat of reaction for the reaction  $c_2h_4$

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web thus for reaction 9.2 in which 1 mole of  $h_2so_4$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$   $\Delta H = -55.8$  kJ as shown note that the enthalpy change  $\Delta H_{products} - \Delta H_{reactants}$  is positive if heat is absorbed that is if  $\Delta H_{products}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove Hess's law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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