



# Tutorial Work From Home Setup

**Raffaela Di Napoli**



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web for example take the reaction of lead ii nitrate with hydrochloric acid to form lead ii chloride and nitric acid shown below  
$$\text{Pb(NO}_3)_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{HNO}_3$$

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total ionic  $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in

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web the net ionic equation for equation 1 is given as equation 3  $\text{s}_2 \text{ aq}$   $\text{fe}_2 \text{ aq}$   $\text{gt}$   $\text{fes s}$  eq 3 8 note that the phase indicator aq is optional on ionic species but is

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web  $\text{cu clo}_4 2$  is soluble all perchlorates are soluble while  $\text{pbso}_4$  is insoluble sulfates are generally soluble except for  $\text{pb}^2$   $\text{ba}^2$   $\text{hg}_2^2$  and  $\text{sr}^2$  the net ionic equation is

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web balanced chemical equation 2  $\text{NaOH (aq)} + \text{MgCl}_2 \text{ (aq)} \rightarrow 2 \text{NaCl (aq)} + \text{Mg(OH)}_2 \text{ (s)}$  full ionic equation 2  $\text{Na}^+ \text{ (aq)} + 2 \text{OH}^- \text{ (aq)} + \text{Mg}^{2+} \text{ (aq)} + 2 \text{Cl}^- \text{ (aq)} \rightarrow 2 \text{Na}^+ \text{ (aq)} + 2 \text{Cl}^- \text{ (aq)} + \text{Mg(OH)}_2 \text{ (s)}$