

Sensory Integration Therapy Techniques

1. Swing Therapy
2. Brushing Protocol
3. Therapeutic Listening
4. Balance Activities
5. Tactile Activities
6. Proprioceptive Activities
7. Visual Motor Activities
8. Sensory Diets
9. Deep Pressure Activities
10. Oral Motor Activities



Therapy Techniques Latest

Joacim Rocklöv



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web the raw score is converted to the standard score using the conversion tables provided and then the standard score is used to calculate the percentile rank for the child s specific

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web in setting out this curve the first curve r 1 is set out in the usual way to point t the theodolite is moved to t and backsighted to t 1 with the horizontal circle reading 180 1 2 set the instrument to read zero and it will then be pointing to t 2
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web compound curve two or more circular arcs of different radii tangent to each other broken back curve combination of a short length of tangent connecting two circular arcs that have centers on the same side

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web using rankine method the curve is set out by the use of deflection angles turned at the point of curvature from the tangent line to the points along the curve this method uses station measurements while two theodolite method does not
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web understand various methods of setting out curves in the field and problems in setting out these curves and know about the need of super elevation and transition curves be uniform and may consist of different gradients for instance rising gradient may be followed by falling gradient and vice versa

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web a compound curve consists of two or more circular curves between two main tangents joined at point of compound curve pcc curve at pc is designated as 1 r 1 l 1 t 1 etc and curve at pt is designated as 2 r 2 l 2 t 2 etc

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web the following are the methods of setting out simple circular curves by linear methods and by the use of chain and tape 1 by ordinates from the long chord 2 by successive bisection of arcs 3 by offsets from the tangents 4 by offsets from chords produced method 1 by ordinates from the long chord fig 11 8

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web 10 2 setting out curves this is the process of establishing the centre line of the curve on the ground by means of pegs at

10 m to 30 m intervals in order to do this the tangent and intersection points must first be fixed in the ground in their correct positions consider figure 10 3

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web the point where the two curves connect namely the point at which the pt of the first curve equals the pc of the second curve is referred to as the point of compound curvature pcc since their tangent lengths vary compound curves fit the topography much better than simple curves

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web oct 6 2022 therefore boyle s law indicates that when the pressure of the object increases the volume of the object decreases and vice versa because volume and pressure s relationship is inversely proportional while the temperature is constant

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web overall the experiment succeeded in showing that temperature and pressure for an ideal gas at constant volume and mass follow the relation of the ideal gas equation differences existed in the experimental graph of temperature versus and pressure and the theoretical curve of temperature versus pressure

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web we then changed the volume between 5ml and 20ml and recorded the pressure at each this part of the lab showed that volume and pressure are inversely proportionate as the pressure would decrease if the volume was increased part b of the experiment looked at the relationship between temperature and pressure

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web procedure 1 the relationship between pressure and volume set the plunger of the syringe to 1 10 the capacity of the

syringe the 20 ml mark if using a 20 ml syringe the volume is read from the edge of the bottom rib of the plunger attach the syringe to the connector at the end of sensor

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1/267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1/273 and concluded that the pressure would approach zero at 1/273 degrees c

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web the volume of 1.00 mol of any gas at stp standard temperature 273.15 k and pressure 1 atm is measured to be 22.414 l we can substitute 101.325 kpa for pressure and 273.15 k for temperature into the ideal gas equation and solve for r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273.15 k and 1 atm 101.325 kpa 1 at stp one mole of an ideal gas has a volume of about 22.4 l this is referred to as the standard molar volume figure 9 18

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web this lab explores the relationships between pressure volume and temperature in an ideal gas the ideal gas law $pV = nRT$ is used and with it we can explore it by seeing how the pressure and volumes changed in the tubes each time that we

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web figure pageindex 10 since the number of moles in a given volume of gas varies with pressure and temperature changes chemists use standard temperature and pressure 273.15 k and 1 atm or 101.325 kpa to report properties of gases

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introduction the ideal gas law is a commonly used formula in chemistry it relates pressure temperature and volume this experiment seeks to prove those relationships by measuring the pressure of gas at various temperatures

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web the pressure volume temperature pvt fluids lab in ku s chemical petroleum engineering department is used to study the properties of fluids under a wide range of pressure and temperature

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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between temperature and pressure the pressure of the 125 ml flask was at 101 36 kpa when the temperature was at 23 2 c

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95 82 kpa 16 0 c 289 k 0 33 97 90 kpa 19 0 c 292 k 0 34 98 97 kpa 99 22 kpa 101 4 kpa 19 1 c

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web the properties of water at constant volume can be represented as a function of pressure and temperature as shown in the diagram below figure 1 the pressure temperature relationship at constant volume the saturation point of water is the condition at which a phase change occurs from liquid to vapour or vapour to liquid

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web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts