

# SOLIDS AND CRYSTAL STRUCTURE

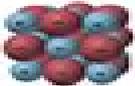
## Solids

- *Molecules are held rigidly in position*
- *Molecules have no freedom of motion*
- *Less empty space in solids than liquids*
- *Virtually incompressible*
- *Molecules are arranged in regular configurations in three dimensions*
- *Has definite shape and volume*

## Intramolecular Forces

- *hold atoms together in a molecule*
- *stabilize individual molecules*
- *involve chemical bonding*

Table 12.2 Comparison of Bonding and Nonbonding (Intermolecular) Forces.

| Force    | Model   | Basis of Attraction               | Energy (kJ/mol) | Example |
|----------|---|-----------------------------------|-----------------|---------|
| Bonding  |   |                                   |                 |         |
| Ionic    |  | Cation-anion                      | 400–4000        | NaCl    |
| Covalent |  | Nuclei-shared e <sup>-</sup> pair | 150–1100        | H–H     |
| Metallic |  | Cations-delocalized electrons     | 75–1000         | Fe      |

## Intermolecular Forces

- *Attractive forces between molecules.*
- *Much weaker than intramolecular forces*

Table 12.2 Comparison of Bonding and Nonbonding (Intermolecular) Forces.

| Nonbonding (Intermolecular) |   |  |         |  |
|-----------------------------|---|--|---------|--|
| Ion-dipole                  |  | Ion-charge-dipole charge                           | 50–500  | $\text{Na}^+ \cdots \text{O} \begin{array}{l} \diagup \\ \diagdown \end{array}$                      |
| H bond                      | $\text{A}-\overset{\delta+}{\text{H}} \cdots \overset{\delta-}{\text{B}}$           | Polar bond to H-dipole charge (high EN of N, O, F) | 10–40   | $\text{O}=\text{C}-\overset{\delta+}{\text{H}} \cdots \overset{\delta-}{\text{O}}-\text{C}=\text{O}$ |
| Dipole-dipole               |  | Dipole charges                                     | 3–25    | $\text{H}-\text{O} \cdots \text{O}-\text{H}$   |
| Ion-induced dipole          |  | Ion-charge-polarizable e <sup>-</sup> cloud        | 3–15    | $\text{Na}^+ \cdots \text{O}_2$  |
| Dipole-induced dipole       |  | Dipole charge-polarizable e <sup>-</sup> cloud     | 2–10    | $\text{H}-\text{O} \cdots \text{O}-\text{H}$   |
| Dispersion (London)         |  | Polarizable e <sup>-</sup> clouds                  | 0.05–40 | $\text{F}-\text{F} \cdots \text{F}-\text{F}$   |

## CATEGORIES OF SOLIDS

### 1. CRYSTALLINE SOLID

- *It possesses rigid and long-range order.*
- *Its atoms, molecules or ions occupy specific positions and have regular structure.*
- *Ex: metals, many ceramics, some polymers.*



### 2. AMORPHOUS SOLID

- *Lacks regular three-dimensional arrangement of atom.*
- *without regular structure*
- *Ex: glass*

# The Crystal Structure Of Solids

**JL Elias**



## The Crystal Structure Of Solids:

**The Crystal Structure of Solids** Penelope Jane Brown, John Bruce Forsyth, 1973 **The Crystal Structure of Solids** Penelope Jane Brown, John Bruce Forsyth, 1973 Structure and Chemistry of Crystalline Solids Bodie Douglas, Shi-Ming Ho, 2007-03-20 Understandable by anyone concerned with crystals or solid state properties dependent on structure Presents a general system using simple notation to reveal similarities and differences among crystal structures More than 300 selected and prepared figures illustrate structures found in thousands of compounds **Crystal Structures M** Ladd, 1999-06-01 This survey of the important types of inorganic and organic crystal structures treats its subject thoroughly and in sufficient depth for undergraduate modules in chemistry courses Features of this book are the instructions for 3D stereoviewing which is central to a full appreciation of the presentation Clear directions for making your own stereo have been provided in the book which enables readers to examine the plentiful stereo of lattices and crystal structures which are illustrated The introductory chapter explains point group and space group symmetry insofar as required to understand lattices and crystal structures Crystal structures are sub divided according to the atomic force mainly responsible for cohesion in the solid state The descriptions of the structures are given in crystallographic terms including data on the space group molecular symmetry and molecular geometry Discussions of bonding theory for each sub division of the structures enhance and strengthen the author's presentation The book stems from the author's successful lecture courses tested and refined in class teaching It draws as necessary on equilibrium thermodynamics and other chemical topics with avoidance of advanced mathematics A level being the prerequisite Examines the important types of inorganic and organic crystal structures Includes instructions for making simple stereoviewers and computer programs Draws as necessary on equilibrium thermodynamics and other chemical topics with avoidance of advanced mathematics *The Structures of Crystals* A.M Glazer, 1987-01-01 The crystalline state is the most stable form of all solids and is very common in nature This book discusses the structure of crystals and the ways in which they are described The basic ideas of symmetry are introduced first The lattice types which are possible are then examined and this leads to the fourteen Bravais lattices There follows a careful treatment of the convolution operation and its applications to crystals and the calculation of Fourier transforms These ideas allow crystal structure to be determined from the results of diffraction experiments The effects of the thermal vibration of atoms and finite crystal size on these experiments is also explained Throughout the text the emphasis is on understanding the results rather than on rigorous proofs Structure of Crystalline Solids, Imperfections and Defects in Crystals Costel Rizescu, Mihaela Rizescu, 2018-03 This book contains a unified approach to crystallography and the structural imperfections and defects found within crystals The reason the authors chose for treating these two subjects together was that the study of the first is necessary for a proper appreciation of the second Understanding the idea of lattice and of the crystal classes firstly will help its relevance to the study of imperfections defects and phase transformations in the second part of the book

Split into two parts the Structure of the Crystalline Solids and the Imperfections and Defects in Crystals and each broken into subchapters this volume combines the classical and exact description of the symmetry of a perfect crystal with the possible geometries of the major defects dislocations stacking faults point line surface and volume defects twins and the effects of martensitic transformation in iron The first part presents a systematic treatment of the basics of crystallography discussing space lattice unit cells symmetry point groups crystal systems in terms of Miller indices for crystallographic points directions and planes and zone axis as well A number of important concepts such as packing factor atomic radius linear planar and volume density polymorphism allotropes interstitial sites in cubic and hexagonal structures and structural features of the most representative compounds used nowadays in a myriad of applications are introduced and carefully explained In the second part the authors guide the reader in a step by step way through point line planar and volume defects with an emphasis on their structural properties A large amount of the latest critically evaluated data for the properties of all elements on the periodic table of the elements including hundreds of up to date crystal structure data are presented Each chapter contains the logical presentation of concepts supported by suitably chosen examples and worked problems Hundreds of illustrations within the text help the reader visualize crystal structures and mathematical objects supporting important topics but also illustrating crystalline structures found in thousands of compounds A concise summary and plenty of review questions and problems at the end reinforce the important key points

#### **Fundamentals of the Physics of Solids** Jenő

Sólyom, 2007-09-19 This book is the first of a three volume series written by the same author It aims to deliver a comprehensive and self contained account of the fundamentals of the physics of solids In the presentation of the properties and experimentally observed phenomena together with the basic concepts and theoretical methods it goes far beyond most classic texts The essential features of various experimental techniques are also explained The text provides material for upper level undergraduate and graduate courses It will also be a valuable reference for researchers in the field of condensed matter physics

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It is to be noted that the purpose of this book is to cover the basic principles and methods of Solid State Physics which are usually included in the course of teaching Physics at the undergraduate graduate and Masters Levels We hope that this book will be useful to the students and teachers in the different universities around the world

**The Pearson Guide To The Scra Examination, 2/E** Thorpe,2008-09

**Mineralogy; an Introduction to the Study of Minerals and Crystals** Edward Henry Kraus,1928

**U.S. Government Research & Development Reports** ,1965

**Journal of the Textile Institute** Textile Institute (Manchester, England),1923 List of members in v 1 8

Chemical News and Journal of Industrial Science ,1911

**Understanding Solids** R. J. D. Tilley,2004-09-03 A modern introduction to the subject taking a unique integrated approach designed to appeal to both science and engineering students Covering a broad spectrum of topics this book includes numerous up to date examples of real materials with relevant applications and a modern treatment of key concepts The science bias allows this book to be equally accessible to engineers chemists and physicists Carefully structured into self contained bite sized chapters to enhance student understanding Questions have been designed to reinforce the concepts presented Includes coverage of radioactivity Reflects a rapidly growing field from the science perspective

*An Introductory Guide to EC Competition Law and Practice* Valentine Korah,1994

The Chemical News and Journal of Industrial Science ,1926

Transactions of the Faraday Society Faraday Society,1918

**Principles of General Chemistry** Stuart Robert Brinkley,1926

**Catalog of the United States Geological Survey Library** ,1974

Solid State Chemistry and Its Applications Anthony R. West,1991-01-08 The first broad account offering a non mathematical unified treatment of solid state chemistry Describes synthetic methods X ray diffraction principles of inorganic crystal structures crystal chemistry and bonding in solids phase diagrams of 1 2 and 3 component systems the electrical magnetic and optical properties of solids three groups of industrially important inorganic solids glass cement and refractories and certain aspects of organic solid state chemistry including the organic metal of new materials

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