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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1 1
 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$
one mole of hcl hcl would be fully neutralized by one mole of naoh naoh

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide naoh and oxalic acid $H_2C_2O_4 \cdot 2H_2O$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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