

Standard Reduction Potentials in Aqueous Solution at 25° C

| Half-reaction | | $E^\circ(\text{V})$ |
|---|--|---------------------|
| $\text{F}_2(\text{g}) + 2 e^-$ | $\rightarrow 2 \text{F}^-$ | 2.87 |
| $\text{Co}^{3+} + e^-$ | $\rightarrow \text{Co}^{2+}$ | 1.82 |
| $\text{Au}^{3+} + 3 e^-$ | $\rightarrow \text{Au}(\text{s})$ | 1.50 |
| $\text{Cl}_2(\text{g}) + 2 e^-$ | $\rightarrow 2 \text{Cl}^-$ | 1.36 |
| $\text{O}_2(\text{g}) + 4 \text{H}^+ + 4 e^-$ | $\rightarrow 2 \text{H}_2\text{O}(\text{l})$ | 1.23 |
| $\text{Br}_2(\text{l}) + 2 e^-$ | $\rightarrow 2 \text{Br}^-$ | 1.07 |
| $2 \text{Hg}^{2+} + 2 e^-$ | $\rightarrow \text{Hg}_2^{2+}$ | 0.92 |
| $\text{Hg}^{2+} + 2 e^-$ | $\rightarrow \text{Hg}(\text{l})$ | 0.85 |
| $\text{Ag}^+ + e^-$ | $\rightarrow \text{Ag}(\text{s})$ | 0.80 |
| $\text{Hg}_2^{2+} + 2 e^-$ | $\rightarrow 2 \text{Hg}(\text{l})$ | 0.79 |
| $\text{Fe}^{3+} + e^-$ | $\rightarrow \text{Fe}^{2+}$ | 0.77 |
| $\text{I}_2(\text{s}) + 2 e^-$ | $\rightarrow 2 \text{I}^-$ | 0.53 |
| $\text{Cu}^+ + e^-$ | $\rightarrow \text{Cu}(\text{s})$ | 0.52 |
| $\text{Cu}^{2+} + 2 e^-$ | $\rightarrow \text{Cu}(\text{s})$ | 0.34 |
| $\text{Cu}^{2+} + e^-$ | $\rightarrow \text{Cu}^+$ | 0.15 |
| $\text{Sn}^{4+} + 2 e^-$ | $\rightarrow \text{Sn}^{2+}$ | 0.15 |

Standard Potentials In Aqueous Solution

Wenbin Ji



Standard Potentials In Aqueous Solution:

Standard Potentials in Aqueous Solution Allen J. Bard, Roger Parsons, Joseph Jordan, 1985-08-27 The best available collection of thermodynamic data The first of its kind in over thirty years this up to date book presents the current knowledge on Standard Potentials in Aqueous Solution Written by leading international experts and initiated by the IUPAC Commissions on Electrochemistry and Electroanalytical Chemistry this remarkable work begins with a thorough review of basic concepts and methods for determining standard electrode potentials Building upon this solid foundation this convenient source proceeds to discuss the various redox couples for every known element The chapters of this practical time saving guide are organized in order of the groups of elements on the periodic table for easy reference to vital material AND each chapter also contains the fundamental chemistry of elements numerous equations of chemical reactions easy to read tables of thermodynamic data and useful oxidation state diagrams Standard Potentials in Aqueous Solution is an ideal handy reference for analytical and physical chemists electrochemists electroanalytical chemists chemical engineers biochemists inorganic and organic chemists and spectroscopists needing information on reactions and thermodynamic data in inorganic chemistry And it is a valuable supplementary text for undergraduate and graduate level chemistry students

Standard Potentials in Aqueous Solution AJ BARD (ED.), 1985 *Environmental Separation of Heavy Metals* Arup K. Sen Gupta, 2001-09-26 This new book explains advanced and emerging technologies for removing heavy metals from wastestreams and contaminated sites Separation processes of this type are critical for meeting stringent regulations of priority pollutants especially arsenic mercury and lead which the text treats in depth After explaining the chemistry of heavy metals and their transport in various media the work offers a comprehensive analysis of strategies for separating metals from groundwater wastewater contaminated soils and industrial sludges Both the basics and the applications of techniques such as ion exchange specialized sorbents novel membranes advanced precipitates and electrokinetic processes are presented with a view to current use and potential for future applications such as resource reuse Information in this volume enables engineers and other investigators to adapt and select the best means to remove and in certain instances recover heavy metals

Standard Electrode Potentials Involving Radicals in Aqueous Solution, 2015 We made recommendations for standard potentials involving select inorganic radicals in aqueous solution at 25 C These recommendations are based on a critical and thorough literature review and also by performing derivations from various literature reports We also summarized the data in tables of standard potentials Gibbs energies of formation radical pKa s and hemocoagulation equilibrium constants In all cases current best estimates of the uncertainties are provided An extensive set of Data Sheets is appended that provide original literature references summarize the experimental results and describe the decisions and procedures leading to each of the recommendations

[Handbook of Electrochemistry](#) Cynthia G. Zoski, 2007-02-07 Electrochemistry plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds biochemical

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Physical Chemistry Peter Atkins, Julio de Paula, 2002 New edition of the overwhelmingly favorite text for the physical chemistry course

Electrochemical Data Dezső Dobos, 1975

Thermodynamics of Dilute Aqueous Solutions Marcel Pourbaix, 1949

I. A Study of the Constitution of Aqueous Solutions of the Hydrogen Soap, Cetyl Sulfonic Acid Robert Case Williams, 1929

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Gmelin Handbook of Inorganic Chemistry Dieter Gras, 1989

The Solubility of Magnetite in Water and in Aqueous Solutions of Acid and Alkali Gerhard Bohnsack, 1987

Gmelin Handbook of Inorganic Chemistry, 1987

Inorganic Chemistry Allen J. Bard, Martin Stratmann, Fritz Scholz, Christopher J. Pickett, 2006-03-31 Electrochemical processes play an increasingly large role in our daily lives whether in producing or saving energy rust protection or nerve stimuli in our bodies This 11 volume encyclopedia provides both an easy introduction to all topics related to modern electrochemistry as well as a comprehensive overview of the subject Unrivalled in its breadth and depth this first class reference work has been created and written by

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Selected Water Resources Abstracts, 1973

Be Beryllium Gudrun Bär, Lieselotte Berg, Gerhard Czack, Dieter Gras, Vera Haase, 1993-10-20 The present Supplement Volume Beryllium A 3 continues and completes the description of the physical properties of the element begun in Supplement Volume A 2 1991 and also treats the electrochemical behavior of the metal The unique combination of the Be properties which was pointed out in Supplement Volume A 2 is also demonstrated in the following chapters of this Volume A 3 13 Electrical Properties 14 Electronic Properties 15 Optical Properties Emission and Impact Phenomena 16 Electrochemical Behavior Starting with the electrical properties Be is a rather good electrical conductor in contrast to what might be expected Superconductivity was studied especially on films Quantum effects which are more pronounced in Be than in most other metals are the reason for numerous investigations of the magnetoresistance and the magnetic breakdown effect The basis for many of the characteristic properties is the unique nature of bonding in Be as a consequence of its peculiar electronic structure and the special shape of its Fermi surface which also gave rise to further numerous studies Detailed cluster calculations were performed to better understand the bonding in the metal Regarding the optical properties the high reflectivity of Be particularly in the infrared region makes it attractive for the fabrication of precision optical surfaces mirrors it is also useful for solar collector surfaces in spacecraft applications Emission and electron and ion impact phenomena as well as neutron optics are also discussed

Properties of Thorium Ions in Solutions Sten Ahrland, Gregory R. Choppin, P.A. Baisden, R.A. Torres, Günter Marx, 1988

British Abstracts, 1929 *British Chemical Abstracts*, 1926

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