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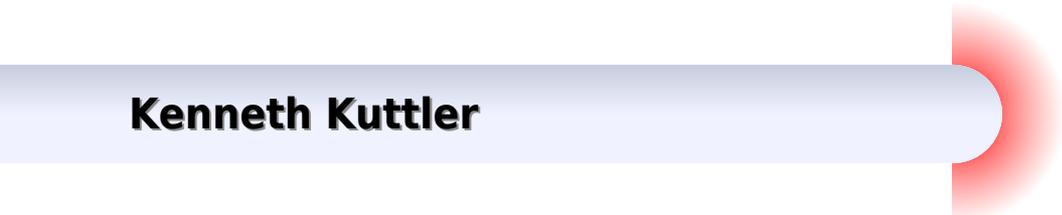
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web correct answer c both a and b q2 the density of 1 m solution of hcl is 1 0585 g ml the molality of the solution is 1 0585 1 0 10 0 0585 correct answer b 1 explanation molality begin array l m left frac 1000 m 1000d mm right end array

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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour pressure and temperature changes concentrations expressed in molality are used

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web 1 molarity mv mass molar mass  $\times 0 4500 \text{ l } 0 825 \text{ g } 141 9579 \text{ g mol } \times 0 0129 \text{ m } 2 \text{ molality } 0 825 \text{ g } 141 9579 \text{ g mol } 0 00581158 \text{ mol } 0 00581158 \text{ mol } 0 4500 \text{ kg } 0 0129 \text{ m } 3 \text{ mole fraction na } 2 \text{ hpo } 4 0 825 \text{ g } 141 9579 \text{ g mol } 0 00581158 \text{ mol h } 2 \text{ o } 450 0 \text{ g } 18 015 \text{ g mol } 24 97918401 \text{ mol}$

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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20  $\text{M}$  3 c a solution with molarity 2 requires 2  $\text{M}$  of  $\text{NaOH}$  per liter so 4  $\times$  2 8  $\text{M}$  4 a a solution of molarity 1 5  $\text{M}$  requires 1 5 mol of  $\text{NaOH}$  to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture  $\text{m}_{\text{mix}}$  can be calculated using the following formula  $\text{m}_{\text{mix}} = \frac{m_1 v_1 + m_2 v_2 + \dots + m_n v_n}{V_{\text{total}}}$  1 2  $\times$  520 480 520 720 624 1000 1 344  $\text{M}$  q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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molarity of a solution containing 325 g of nacl dissolved in 750 ml of solution 3 140 g of kcl is dissolved in 600 ml of water  
what is the molarity 4 724 4 g of ammonium phosphate in 4500 ml of alcohol

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