



CHRISTMAS PLAY PROGRAM

DECEMBER 12, 2050
TOWN THEATER

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- 
-
- 6:00 PM** Welcome and Introduction
 - 6:15 PM** Performance Act 1: "The Christmas Story Begins"
 - 6:45 PM** Intermission and Refreshments
 - 7:00 PM** Performance Act 2: "The Journey to Bethlehem"
 - 7:30 PM** Performance Act 3: "The Arrival of the Magi"
 - 8:00 PM** Special Musical Interlude by the Choir
 - 8:15 PM** Audience Participation: Christmas Carol Sing-Along
 - 8:30 PM** Performance Finale: "The True Meaning of Christmas"
 - 8:45 PM** Closing Remarks and Thank You

Standard Christmas Program 47

M Mark



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Catalog of Copyright Entries. Third Series Library of Congress. Copyright Office, 1955 Catalog of Copyright Entries
Library of Congress. Copyright Office, 1970 **The National Union Catalog, Pre-1956 Imprints** Library of
Congress, American Library Association. Committee on Resources of American Libraries. National Union Catalog
Subcommittee, 1968 **Home Mission Monthly**, 1909 **Paperbound Books in Print**, 1991 **Handbook for**
Guardians of Camp Fire Girls Camp Fire Girls, 1930 Handbook for volunteer guardians that is adult leaders of Camp Fire
groups Many of the early traditions of Camp Fire Girls including dresses symbolism and language are culturally appropriated
from Indigenous peoples The contents and design of this volume are representative of this **Tabulation of Voluntary**
Standards and Certification Programs for Consumer Products Sophie J. Chumas, 1973 **Presleyana** Jerry
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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2 \text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ \rightarrow $2 \text{Cr}_2\text{O}_7^{2-} \text{s}$ label 4 2 3

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web jan 29 2020 the ionic form of this reaction is $2 \text{Na}^+ \text{aq}$ $2 \text{Cl}^- \text{aq}$ $\text{Cu}^{2+} \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $2 \text{Na}^+ \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $\text{CuCl}_2 \text{s}$ the sodium ions

and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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 net ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} + \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g} \rightarrow \text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions
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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3\text{PO}_4 \text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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